

Today in Math History...

-18 October 1921

Niels Bohr introduces quantum model of the atom.

The Bohr model of the hydrogen atom, where negatively charged electrons confined to Atomic shells encircle a small positively charged atomic nucleus, and that an electron jump between orbits must be accompanied by an emitted or absorbed amount of electromagnetic energy $h\nu$. The orbits that the electrons travel in are shown as grey circles; their radius increases n^2 , where n is the principal quantum number. The $3 \rightarrow 2$ transition depicted here produces the first line of the Balmer series, and for hydrogen ($Z = 1$) results in a photon of wavelength 656 nm (red).

